JEE-MAIN – JUNE, 2022

(Held On Tuesday 27th June, 2022) **TIME : 3 : 00 PM to 6 : 00 PM**

Chemistry

Test Pattern : JEE-MAIN Maximum Marks : 120

Topic Covered: FULL SYLLABUS

Important instruction:

1. Use Blue / Black Ball point pen only.
2. There are three sections of equal weightage in the question paper **Physics, Chemistry** and **Mathematics** having 30 questions in each subject. Each paper have 2 sections A and B.
3. You are awarded +4 marks for each correct answer and –1 marks for each incorrect answer.
4. Use of calculator and other electronic devices is not allowed during the exam.
5. No extra sheets will be provided for any kind of work.

Name of the Candidate (in Capitals) _____

Father's Name (in Capitals) _____

Form Number : in figures _____

: in words _____

Centre of Examination (in Capitals): _____

Candidate's Signature: _____ Invigilator's Signature : _____

Rough Space

YOUR TARGET IS TO SECURE GOOD RANK IN JEE-MAIN

FINAL JEE–MAIN EXAMINATION – JUNE, 2022

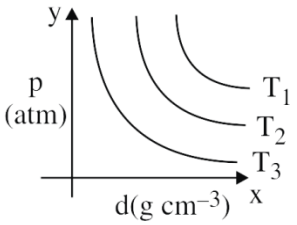
(Held On Monday 27th June, 2022) **TIME : 3 : 00 PM to 6 : 00 PM**

CHEMISTRY — TEST PAPER WITH SOLUTIONS

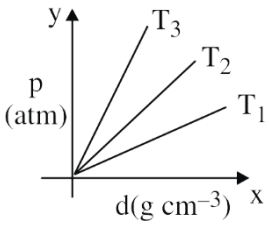
SECTION-A

1. Which amongst the given plots is the correct plot for pressure (p) vs density (d) for an ideal gas ?

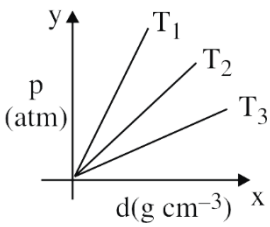
(A) $T_3 > T_2 > T_1$



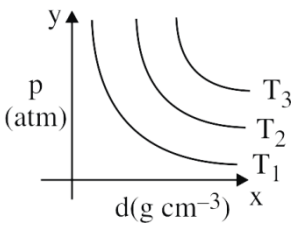
(B) $T_3 > T_2 > T_1$



(C) $T_3 > T_2 > T_1$



(D) $T_3 > T_2 > T_1$

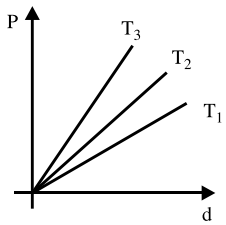


Official Ans. by NTA (B)

Allen Ans. (B)

Sol. P vs d :

$$P = \left(\frac{RT}{M}\right) d$$



$T_3 > T_2 > T_1$

2. Identify the **incorrect** statement for PCl_5 from the following.

(A) In this molecule, orbitals of phosphorous are assumed to undergo sp^3d hybridization.

(B) The geometry of PCl_5 is trigonal bipyramidal.

(C) PCl_5 has two axial bonds stronger than three equatorial bonds.

(D) The three equatorial bonds of PCl_5 lie in a plane.

Official Ans. by NTA (C)

Allen Ans. (C)

Sol. In PCl_5, axial bonds are weaker than equatorial.

3. **Statement I : Leaching of gold** with cyanide ion in absence of air / O_2 leads to cyano complex of Au(III).

Statement II : Zinc is oxidized during the displacement reaction carried out for gold extraction.

In the light of the above statements, choose the correct answer from the options given below.

(A) Both Statement I and Statement II are correct

(B) Both Statement I and Statement II are incorrect

(C) Statement I is correct but Statement II is incorrect

(D) Statement I is incorrect but Statement II is correct

Official Ans. by NTA (D)

Allen Ans. (D)

Sol. Statement-1 : wrong, Au^+ is correct, not Au^{+3}

Statement-2 : correct

4. The correct order of increasing intermolecular hydrogen bond strength is

(A) $HCN < H_2O < NH_3$

(B) $HCN < CH_4 < NH_3$

(C) $CH_4 < HCN < NH_3$

(D) $CH_4 < NH_3 < HCN$

Official Ans. by NTA (C)

Allen Ans. (C)

Sol. Order of H-Bonding

$CH_4 < HCN < NH_3$

NCH . . . NCH

H_2NH . . . NH_3

5. The correct order of increasing ionic radii is

(A) $Mg^{2+} < Na^{+} < F^{-} < O^{2-} < N^{3-}$

(B) $N^{3-} < O^{2-} < F^{-} < Na^{+} < Mg^{2+}$

(C) $F^{-} < Na^{+} < O^{2-} < Mg^{2+} < N^{3-}$

(D) $Na^{+} < F^{-} < Mg^{2+} < O^{2-} < N^{3-}$

Official Ans. by NTA (A)

Allen Ans. (A)

Sol. $N^{-3} > O^{-2} > F^{-} > Na^{+} > Mg^{+2}$ (Radii)

(Isoelectronic species)

6. The gas produced by treating an aqueous solution of ammonium chloride with sodium nitrite is

(A) NH_3 (B) N_2

(C) N_2O (D) Cl_2

Official Ans. by NTA (B)

Allen Ans. (B)

Sol. $NH_4Cl + NaNO_2 \rightarrow NH_4NO_2 + NaCl$

$\downarrow$

$N_2 + 2H_2O$

7. Given below are two statements: one is labelled as Assertion A and the other is labelled as Reason R.

Assertion A : Flourine forms one oxoacid.

Reason R : Flourine has smallest size amongst all halogens and is highly electronegative

In the light of the above statements, choose the most appropriate answer from the options given below.

(A) Both A and R are correct and R is the correct explanation of A.

(B) Both A and R are correct but R is NOT the correct explanation of A.

(C) A is correct but R is not correct.

(D) A is not correct but R is correct

Official Ans. by NTA (A)

Allen Ans. (A)

Sol. Both A and R are correct and R is the correct explanation of A.

8. In 3d series, the metal having the highest M^{2+}/M standard electrode potential is

(A) Cr (B) Fe

(C) Cu (D) Zn

Official Ans. by NTA (C)

Allen Ans. (C)

Sol. $Cr^{+2}/Cr \rightarrow -0.90$ V

$Fe^{+2}/Fe \rightarrow -0.44$ V

$Cu^{+2}/Cu \rightarrow +0.34$ V

$Zn^{+2}/Zn \rightarrow -0.76$ V

So Ans. Cu^{+2}/Cu

9. The 'f' orbitals are half and completely filled, respectively in lanthanide ions

(Given: Atomic no. Eu, 63; Sm, 62; Tm, 69; Tb, 65; Yb, 70; Dy, 66]

(A) Eu^{2+} and Tm^{2+} (B) Sm^{2+} and Tm^{3+}

(C) Tb^{4+} and Yb^{2+} (D) Dy^{3+} and Yb^{3+}

Official Ans. by NTA (C)

Allen Ans. (C)

Sol. $Tb \rightarrow 4f^9 6s^2$

$Tb^{+4} \rightarrow 4f^7$

$Yb \rightarrow 4f^{14} 6s^2$

$Yb^{+2} \rightarrow 4f^{14}$

10. Arrange the following coordination compounds in the increasing order of magnetic moments. (Atomic numbers: Mn = 25; Fe = 26)

(A) $[FeF_6]^{3-}$

(B) $[Fe(CN)_6]^{3-}$

(C) $[MnCl_6]^{3-}$ (high spin)

(D) $[Mn(CN)_6]^{3-}$

(A) A < B < D < C (B) B < D < C < A

(C) A < C < D < B (D) B < D < A < C

Official Ans. by NTA (B)

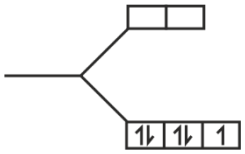
Allen Ans. (B)

Sol. (A) $[FeF_6]^{3-}$

$Fe^{+3} \rightarrow 3d^5\ 4s^0$

n = 5

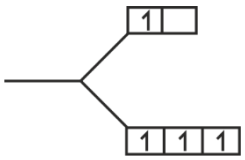
(B) $[Fe(CN)_6]^{3-}$



$Fe^{+3} \rightarrow 3d^5\ 4s^0$

n = 1

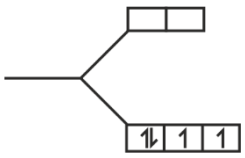
(C) $[MnCl_6]^{3-}$



$Mn^{+3} \rightarrow 3d^4\ 4s^0$

n = 4

(D) $[Mn(CN)_6]^{3-}$



$Mn^{+3} \rightarrow 3d^4\ 4s^0$

n = 2

$\mu \Rightarrow A > C > D > B$

11. On the surface of polar stratospheric clouds, hydrolysis of chlorine nitrate gives A and B while its reaction with HCl produces B and C. A, B and C are, respectively

(A) HOCl, HNO_3, Cl_2

(B) Cl_2, HNO_3, HOCl

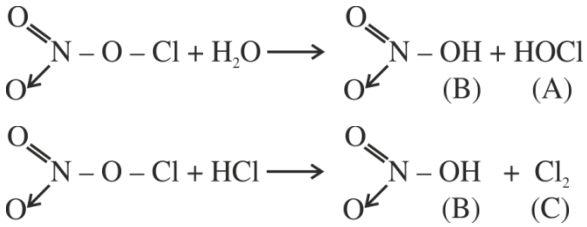
(C) $HClO_2$, HNO_2, HOCl

(D) HOCl, HNO_2, Cl_2O

Official Ans. by NTA (A)

Allen Ans. (A)

Sol.



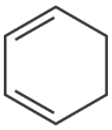
12. Which of the following is most stable?

(A)

(B)

(C)

(D)



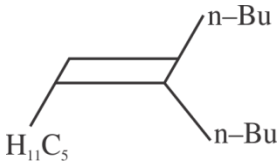
Official Ans. by NTA (A)

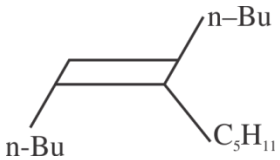
Allen Ans. (A)

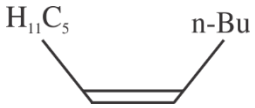
Sol.  is most stable as it is aromatic.

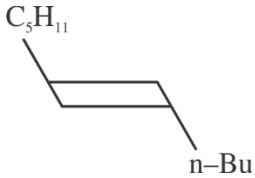
13. What will be the major product of following sequence of reactions?

$$n-Bu-\equiv \xrightarrow[\text{(ii) Lindlar cat, } H_2]{\text{(i) } n-BuLi,\ n-C_5H_{11}Cl}$$

(A) 

(B) 

(C) 

(D) 

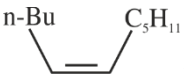
Official Ans. by NTA (C)

Allen Ans. (C)

Sol. $n-Bu-C\equiv CH \xrightarrow[\text{(acid base reaction)}]{nBuLi}$

$n-Bu-C\equiv C^{-}Li^{+}$

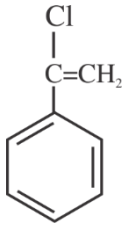
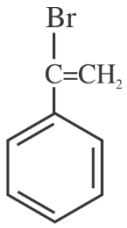
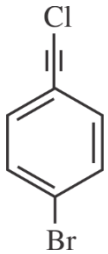
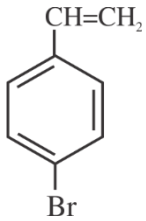
$n-C_5H_{11}Cl \downarrow$ (SN reaction)



$\xleftarrow[\text{Lindlar's Catalyst}]{H_2} n-Bu-C\equiv C-C_5H_{11}$

14. Product 'A' of following sequence of reactions is

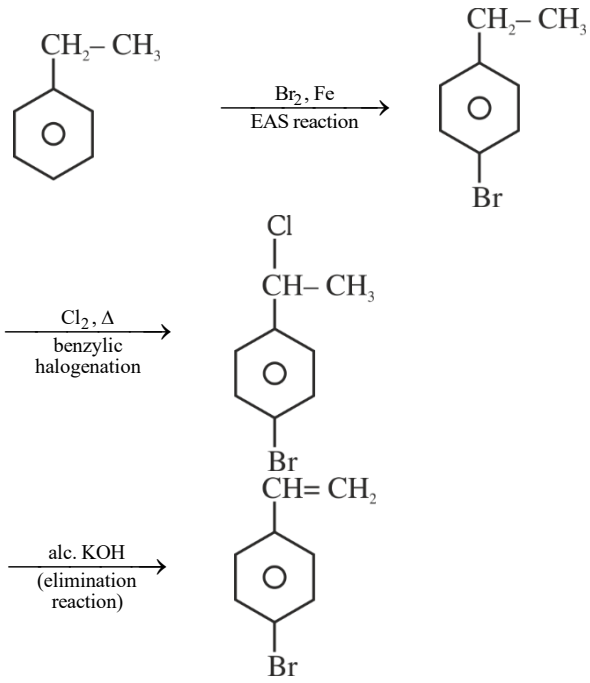
$$\text{Ethylbenzene} \xrightarrow[\text{(b) } Cl_2,\Delta \text{ (c) alc. KOH}]{\text{(a) } Br_2.Fe} \text{'A'(Major product)}$$

(A)	
(B)	
(C)	
(D)	

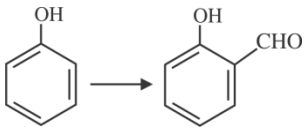
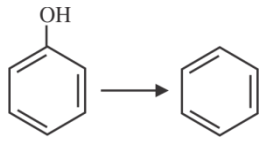
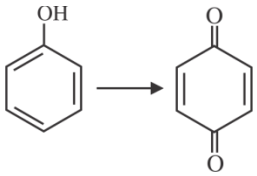
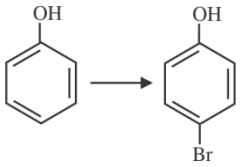
Official Ans. by NTA (D)

Allen Ans. (D)

Sol.



15. Match List I with List II

List I	List II
A. 	I. Br_2 in CS_2
B. 	II. $Na_2Cr_2O_7/H_2SO_4$
C. 	III. Zn
D. 	IV. $CHCl_3$/NaOH

Choose the correct answer from the options given below:

(A) A-IV, B-III, C-II, D-I

(B) A-IV, B-III, C-I, D-II

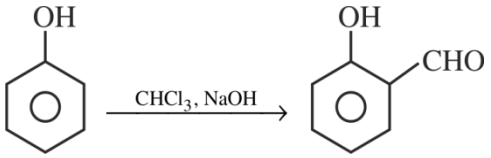
(C) A-II, B-III, C-I, D-IV

(D) A-IV, B-II, C-III, D-I

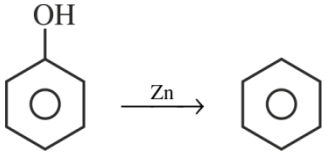
Official Ans. by NTA (A)

Allen Ans. (A)

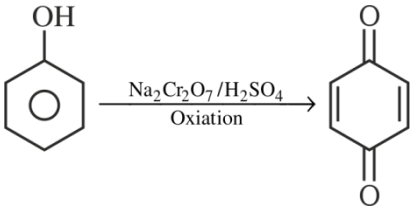
Sol. (A)



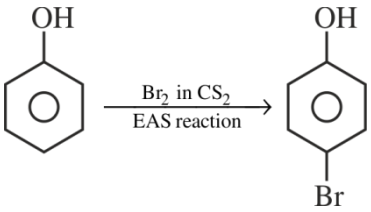
(B)



(C)



(D)



16. Decarboxylation of all six possible forms of diaminobenzoic acids $C_6H_3(NH_2)_2COOH$ yields three products A, B and C. Three acids give a product 'A', two acids gives a product 'B' and one acid give a product 'C'. The melting point of product 'C' is

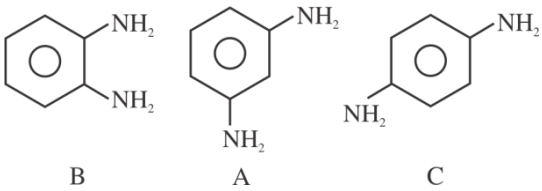
(A) 63ºC (B) 90ºC

(C) 104ºC (D) 142ºC

Official Ans. by NTA (D)

Allen Ans. (D)

Sol.



M.P. 142ºC

17. Which is true about Buna-N?

(A) It is a linear polymer of 1, 3-butadiene.

(B) It is obtained by copolymerization of 1, 3-butadiene and styrene.

(C) It is obtained by copolymerization of 1, 3-butadiene and acrylonitrile.

(D) The suffix N in Buna-N stands for its natural occurrence

Official Ans. by NTA (C)

Allen Ans. (C)

Sol. It is copolymerization of 1, 3-butadiene and acrylonitrile.

18. Given below are two statements.

Statments I: Maltose has two α-D-glucose units linked at C_1 and C_4 and is a reducing sugar.

Statement II: Maltose has two monosaccharides: α-D-glucose and β-D-glucose linked at C_1 and C_6 and it is a non-reducing sugar.

In the light of the above statements, choose the correct answer from the options given below.

(A) Both Statement I and Statement II are true

(B) Both Statement I and Statement II are false

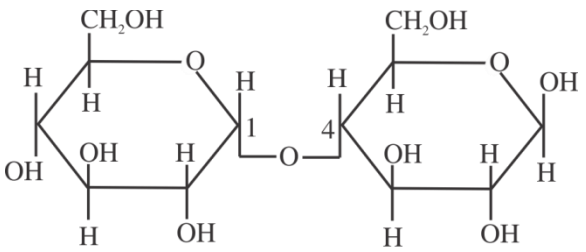
(C) Statement I is true but Statement II is false

(D) Statement I is false but Statement II is true

Official Ans. by NTA (C)

Allen Ans. (C)

Sol.



Maltose

19. Match List I with List Ii

List I	List II
A. Antipyretic	I. Reduces pain
B. Analgesic	II. Reduces stress
C. Tranquilizer	III. Reduces fever
D. Antacid	IV. Reduces acidity (Stomach)

Choose the correct answer from the options given below:

(A) A-III, B-I, C-II, D-IV

(B) A-III, B-I, C-IV, D-II

(C) A-I, B-IV, C-II, D-III

(D) A-I, B-III, C-II, D-IV

Official Ans. by NTA (A)

Allen Ans. (A)

Sol.

A. Antipyretic	Reduces fever
B. Analgesic	Reduces pain
C. Tranquilizer	Reduces stress
D. Antacid	Reduces acidity (Stomach)

20. Match List I with List II

List I (Anion)	List II (Gas evolved on reaction with dil. H_2SO_4)
A. CO_3^{2-}	I. Colourless gas which turns lead acetate paper black
B. S^{2-}	II. Colourless gas which turns acidified potassium dichromate solution green.
C. SO_3^{2-}	III. Brown fumes which turns acidified KI solution containing starch blue.
D. NO_2^-	IV. Colourless gas evolved with brisk effervescence, which turns lime water milky.

Choose the correct answer from the options given below:

(A) A-III, B-I, C-II, D-IV

(B) A-II, B-I, C-IV, D-III

(C) A-IV, B-I, C-III, D-II

(D) A-IV, B-I, C-II, D-III

Official Ans. by NTA (D)

Allen Ans. (D)

Sol. CO_3^{2-} will give CO_2(g) which will turns lime water milky.

S^{2-} will give H_2S (g), will turns lead acetate paper black

SO_3^{2-} will give SO_2 (g), which will turns acidified potassium dichromate solution green.

NO_2^- will give brown NO_2(g) will turn KI solution blue.

SECTION-B

1. 116 g of a substance upon dissociation reaction, yields 7.5 g of hydrogen, 60g of oxygen and 48.5 g of carbon. Given that the atomic masses of H, O and C are 1, 16 and 12 respectively. The data agrees with how many formulae of the following?

(A) CH_3COOH (B) HCHO

(C) CH_3OOCH_3 (D) CH_3CHO

Official Ans. by NTA (2)

Allen Ans. (2)

Sol. $\%H = \frac{7.5}{116} \times 100 = 6.5$

$\%O = \frac{60}{116} \times 100 = 51.7$

$\%C = \frac{48.5}{116} \times 100 = 41.8$

Relative atomicities = $H \Rightarrow 6.5$

$O \Rightarrow \frac{51.7}{16} = 3.25$

$C \Rightarrow \frac{41.8}{12} = 3.5$

Emperically formula is approx.. CH_2O

(A) $C_2H_4O_2$ (B) CH_2O relate to this formula.

2. Consider the following set of quantum numbers

	n	l	m_l
A.	3	3	-3
B.	3	2	-2
C.	2	1	+1
D.	2	2	+2

The number of correct sets of quantum numbers is _____

Official Ans. by NTA (2)

Allen Ans. (2)

Sol. Quantum no. of set (B) and (C) can be correct.

(A) and (D) are wrong as $n = \ell$ is not possible.

3. BeO reacts with HF in presence of ammonia to give [A] which on thermal decomposition produces [B] and ammonium fluoride. Oxidation state of Be in [A] is_____

Official Ans. by NTA (2)

Allen Ans. (2)

Sol.

$$BeO + HF + NH_3 \rightarrow \underset{[A]}{(NH_4)_2\ [BeF_4]} \xrightarrow{\Delta} \underset{[B]}{NH_4F + BeF_2}$$

Oxidation state of Be in A is (+2)

4. When 5 moles of He gas expand isothermally and reversibly at 300 K from 10 litre to 20 litre, the magnitude of the maximum work obtained is ____ J. [nearest integer] (Given: R = 8.3 J $K^{-1}mol^{-1}$ and log 2 = 0.3010)

Official Ans. by NTA (8630)

Allen Ans. (8630)

Sol. n = 5 mol

T = 300 K

V_1 = 10 L

V_2 = 20 L

$w = -nRT\,\ell n\frac{V_2}{V_1}$

$= -5 \times 8.3 \times 300 \times \ell n\frac{20}{10}$

= –8630.38 J

5. A solution containing 2.5×10^{-3} kg of a solute dissolved in 75×10^{-3} kg of water boils at 373.535 K. The molar mass of the solute is _____ g mol^{-1}. [nearest integer] (Given: K_b (H_2O) = 0.52 K Kg mol^{-1} , boiling point of water = 373.15K)

Official Ans. by NTA (45)

Allen Ans. (45)

Sol. w = 2.5 g $\quad$ $K_b = 0.52$

$w_{solvent} = 75$ g $\quad$ M = Mol. Wt. of solute

$T'_B = 373.535\,K$

$T_B^o = 373.15\,K$

$\Delta T_B = 0.385 = K_b$ molality

$$0.385 = 0.52 \times \left(\frac{2.5}{M} \times \frac{1000}{75}\right)$$

$M = 45$ g mol^{-1}

6. pH value of 0.001 M NaOH solution is_____.

Official Ans. by NTA (11)

Allen Ans. (11)

Sol. 0.001 M NaOH

$[OH^-] = 10^{-3}$

pOH = 3

pH = 11

7. For the reaction taking place in the cell:

$Pt(s) \mid H_2(g) \mid H^+(aq) \parallel Ag^+(aq) \mid Ag(s)$

$E^o_{Cell} = +0.5332$ V.

The value of $\Delta_f G^0$ is _____ kJ mol^{-1}. (in nearest integer)

Official Ans. by NTA (51)

Allen Ans. (51 or 103)

Sol. $\frac{1}{2}H_2 + Ag^+ \rightarrow H^+ + Ag$

$\Delta G° = -nE° F$

$= -1 \times 0.5332 \times 96500$ J

$= -51.35$ kJ

(n = 2 for $H_2 + 2Ag^+ \rightarrow 2H^+ + 2Ag$)

8. It has been found that for a chemical reaction with rise in temperature by 9K the rate constant gets doubled. Assuming a reaction to be occurring at 300 K, the value of activation energy is found to be _____ kJ mol^{-1}. [nearest integer]

(Given ln 10 = 2.3, R = 8.3 JK^{-1}mol^{-1}, log2 = 0.30)

Official Ans. by NTA (59)

Allen Ans. (59)

Sol. $$\log_{10}\frac{K_2}{K_1} = \frac{E_a}{2.303R}\left(\frac{1}{300} - \frac{1}{309}\right)$$

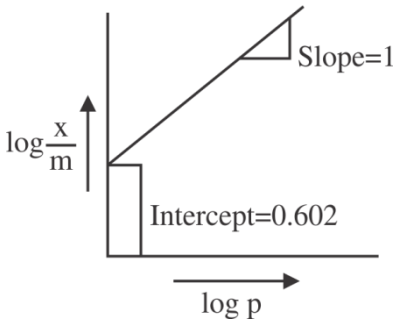
$$0.3 = \frac{E_a}{2.303 \times 8.3}\left(\frac{9}{300 \times 309}\right)$$

$$E_a = \frac{0.3 \times 2.303 \times 8.3 \times 300 \times 309}{9}$$

= 59065.04 J

$E_a = 59.06$ kJ

9.



If the initial pressure of a gas is 0.03 atm, the mass of the gas adsorbed per gram of the adsorbent is _____ × 10^{-2}g.

Official Ans. by NTA (12)

Allen Ans. (12)

Sol. $$\frac{x}{m} = kP^{\frac{1}{n}}$$

$$\log\frac{x}{m} = \log k + \frac{1}{n}\log P$$

From graph

$$\text{Slope} = \frac{1}{n} = 1 \Rightarrow n = 1$$

Intercept = log k = 0.602

k = 4

$$\frac{x}{m} = 4 \times (0.03)^{\frac{1}{1}}$$

$$\frac{x}{m} = 12 \times 10^{-2}$$

10. 0.25 g of an organic compound containing chlorine gave 0.40 g of silver chloride in Carius estimation. The percentage of chlorine present in the compound is _____. [in nearest integer]

(Given: Molar mass of Ag is 108 g mol^{-1} and that of Cl is 35.5 g mol^{-1})

Official Ans. by NTA (40)

Allen Ans. (40)

Sol. wt. of organic compound = 0.25 g

$$\text{mass of Cl} = \frac{35.5}{143.5} \times 0.4\text{g}$$

mass % of Cl in the organic compound

$$= \frac{35.5 \times 0.4}{143.5 \times 0.25} \times 100$$

= 39.58%